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Chemistry - Chapter 9 Assessment Answers 9.1 Assessment 3. The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8. 4. from the number of electrons lost 5. -ite or -ate 6. Group 1A metals, 1+; Group 3A (aluminum), 3+ Group 5A nonmetals, 3-.

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Solutions Manual Chemistry: Matter and Change • Chapter 9 143 CHAPTER 9 SOLUTIONS MANUAL Problem-Solving Lab page 294 Halogen Atomic Radius (pm) Ionic Energy (kJ/mol) Electronegativity Fluorine 72 1681 3.98 Chlorine 100 1251 3.16 Bromine 114 1140 2.96 Iodine 133 1008 2.66 Astatine 140 920 2.2 Properties of Halogens 1.

Chemical Reactions

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Chapter 9 Practice Test - Naming and Writing Chemical Formulas Answer Section MATCHING 1. ANS: A PTS: 1 DIF: L1 REF: p. 253 OBJ: 9.1.1 Identify the charges of monatomic ions by using the periodic table, and name the ions. STA: Ch.5.a 2.ANS: H PTS: 1 DIF: L1 REF: p. 254 OBJ: 9 ...

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e. molar mass of $\text{Fe}(\text{NO}_3)_3 = 241.9 \text{ g}$ 272 g $\text{Fe}(\text{NO}_3)_3$ 1 mol 241.9 g = 1.12 mol $\text{Fe}(\text{NO}_3)_3$ f. molar mass of $\text{FeSO}_4 = 151.9 \text{ g}$ 12.7 mg FeSO_4 1 g 1000 mg 1 mol 151.9 g = 8.36 10^{-5} mol FeSO_4 4 g. molar mass of $\text{LiOH} = 23.95 \text{ g}$ 6.91 103 g LiOH 1 mol 23.95 g = 288.5 = 289 mol LiOH 14. a. molar mass of $\text{CaCO}_3 = 100.1 \text{ g}$ 2.21 10^{-4} mol CaCO_3 100.1 g 1 mol = 0 ...

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Chemistry: Matter and Change Chapter 5 c. d. Name CHAPTER Applying Scientific Methods Date Class CHAPTER ASSESSMENT A chemist isolated four samples, A, B, C, and D. She obtained the following atomic emission spectra of the samples. 400 500 600 700 29 Nanometers 1.

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